

Chemical Equations

Isaac_Toms Brainy's

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1 Questions

Balance the following chemical equations:

1. $\text{N}_2 + \text{O}_2 \longrightarrow \text{N}_2\text{O}_5$
2. $\text{MoS}_2(\text{s}) + \text{O}_2(\text{g}) \longrightarrow \text{MoO}_3(\text{s}) + \text{SO}_2(\text{g})$
3. $\text{FeO}(\text{s}) + \text{O}_2(\text{g}) \longrightarrow \text{S Fe}_2\text{O}_3(\text{s})$
4. $\text{Mg}_3\text{Si}_2\text{O}_5(\text{OH})_4 + \text{CO}_2 \longrightarrow \text{MgCO}_3 + \text{SiO}_2 + \text{H}_2\text{O}$
5. $\text{AsCl}_3 + \text{NaBH}_4 \longrightarrow \text{AsH}_3 + \text{NaCl} + \text{BCl}_3$
6. $3\text{CuS} + \text{HNO}_3 \longrightarrow \text{CuSO}_4 + \text{NO} + \text{H}_2\text{O}$
7. $\text{C}_4\text{H}_{10} + \text{F}_2 \longrightarrow \text{C}_4\text{F}_{10} + \text{HF}$
8. $\text{FeS}_2 + \text{O}_2 \longrightarrow \text{Fe}_2\text{O}_3 + \text{SO}_2$
9. $\text{NO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \longrightarrow \text{HNO}_3(\text{aq}) + \text{NO}(\text{g})$
10. $\text{NH}_3(\text{g}) + \text{O}_2(\text{g}) \longrightarrow \text{N}_2\text{O}(\text{g}) + \text{H}_2\text{O}(\text{l})$
11. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \text{O}_2 \longrightarrow \text{C}_2\text{H}_2(\text{CO})_2\text{O} + \text{H}_2\text{O}$
12. $(\text{NH}_4)_2\text{Cr}_2\text{O}_7(\text{s}) + \text{Cr}_2\text{O}_3(\text{s}) \longrightarrow \text{N}_2(\text{g}) + \text{H}_2\text{O}(\text{g})$
13. $\text{Eu}(\text{s}) + \text{HF}(\text{g}) \longrightarrow \text{EuF}_3(\text{s}) + \text{H}_2(\text{g})$
14. $\text{Ca}_3(\text{PO}_4)_2(\text{s}) + \text{H}_2\text{SO}_4(\text{aq}) \longrightarrow \text{CaSO}_4(\text{s}) + \text{H}_3\text{PO}_4(\text{aq})$
15. $\text{B}_2\text{O}_3(\text{s}) + \text{C}(\text{s}) + \text{Cl}_2(\text{g}) \longrightarrow \text{BCl}_3(\text{g}) + \text{CO}(\text{g})$
16. $\text{CuFeS}_2 + \text{SiO}_2 + \text{O}_2 \longrightarrow \text{Cu}_2\text{S} + \text{FeSiO}_3 + \text{SO}_2$
17. $\text{I}_2 + 10\text{HNO}_3 \longrightarrow \text{HIO}_3 + \text{NO}_2 + \text{H}_2\text{O}$
18. $\text{Al}(\text{s}) + \text{NH}_4\text{ClO}_4(\text{s}) \longrightarrow \text{Al}_2\text{O}_3(\text{s}) + \text{AlCl}_3(\text{s}) + \text{NO}(\text{g}) + \text{H}_2\text{O}(\text{g})$
19. $\text{Ca}_5\text{F}(\text{PO}_4)_3 + 9\text{SiO}_2 + \text{C} \longrightarrow 9\text{CaSiO}_3 + \text{CaF}_2 + \text{CO} + \text{P}$
20. $\text{FeCr}_2\text{O}_4 + \text{Na}_2\text{CO}_3 + \text{O}_2 \longrightarrow \text{Na}_2\text{CrO}_4 + 2\text{Fe}_2\text{O}_3 + 8\text{CO}_2$